

CALIFORNIA STATE SCIENCE FAIR 2014 PROJECT SUMMARY

Name(s) **Project Number** Jonah P. Tenenbaum 34570 **Project Title How Low Can You Go? Abstract Objectives/Goals** The objective is to analyze why salt is used to melt ice on roads and determine t lowers the freezing point of water and to prove the hypothesis: If water is mixed with salt, will freeze at a lower temperature than water without salt. Methods/Materials ferent concentrations of sea salt. Three plastic cups were filled with 8 ounces of distilled water and diff control held only one cup of distilled water, while the other two cups had 1 of 2 tbsps of salt dissolved into them. All three cups were placed in the freezer and then the temperatures of each cup were measured by a food thermometer and recorded at timed intervals of 20 minutes and ice formed and the temperatures were stable. The data from three trials were averaged. Results The cup with 2 tbsp of salt froze at the lowest temperature with an average 14.7°F, the water with 1 tbsp of salt froze at an average temperature of 22.5°, and fresh water finze at an average temperature of 30.6°F. As shown by the data, the water with 2 tbsp of sal froze at the lowest temperature, the water with 1 tbsp of salt froze in the middle, and the fresh water froze at the highest temperature. Conclusions/Discussion According to the data collected, the hypothesis appears to be supported. The water did freeze at a lower temperature when salt was added to it in all three trials. The experiment supports the real life application of using salt to prevent and/or clear ice or roads, because salt lowers the freezing point of ice below the temperature of the surrounding are Summary Statement why salt melts the ice on roads and/or prevents it from forming by lowering the freezing temperature of the water. **Help Received** My parents bought supplies, revised the document and the board. Mr. Joseph (my science teacher) gave me ideas for edits.