

CALIFORNIA STATE SCIENCE FAIR 2014 PROJECT SUMMARY

Name(s) **Project Number** Joshua Batres 34598 **Project Title** Coin Batteries: Which Is a Better Electrolyte Solution, Acid or Base, According to Their Voltages? What Are Their Volta **Abstract Objectives/Goals** This project is about which electrolyte solution is better, acid or base. I used an exam e of acid which is vinegar and an example of base which is baking soda. Then I made a coin battery from an English penny and a piece of zinc. Methods/Materials Materials: an English coin, pieces of tissue paper, vinegar, baking soda, tap water, Nothespin(paperclip), a PH meter, a multimeter Procedure: 1.Dissolve 2g of baking soda into 1/3 cup of water. Take 1/2 cup of vinezar into a cup. Cut two pieces of tissue paper and wet one of them with vinegar and another one with balang soda solution. 2:Measure the PH of the liquids. Place the papers between an English penny and a piece of zinc. Hold them with a clothespin or a paperclip. Connect wires the metals and measure electricity of the battery. 3:Change the piece of tissue to another one. Before change it, wash the metals with water. 4: Repeat this procedure 3 times. Results average PH of the liquids: vinegar: 2.2 baking soda: 7.7 average voltages of the battery vinegar batteries: 0.98v baking soda batteries: 0.83v **Conclusions/Discussion** Vinegar is an acid and baking soda is a base. Results show that an acidic electrolyte solution gives me a higher voltages than a basic electronic solution Both of them worked as sectrolyte solution of batteries. Because acidic electrolyte solution gave me a higher voltage than baking soda, I can say that an acidic electrolyte solution is better than a basic electrolyte solution. **Summary Statement** electrolyte solutions of batteries. This project is about Help Received My brother helped me to wash my batterie's metals.